**QUESTIONS**

**TEXTBOOK**

***Page(s)* ANSWERS**

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| Describe the liquid state according to the kinetic-molecular theory. | 333 | Liquids have the medium energy level of the three state. They have enough energy for the molecules to break free from their solid lattice but not enough to separate from being in groups of molecules, they move around slowly and randomly in groups, with a moderate level of kinetic energy to power them. |
| list the properties of liquids. | 333-334 | 1. take the shape of its container  2. some space between particles  3. particles move at a medium speed sometimes bouncing ioff each other  4. has a fixed volume |
| explain why liquids in a test tube form a meniscus. | 334 | Formation of a meniscus is a surface tension effect. Depending on the chemical properties of the various components, there may be a stronger affinity between the liquid and the solid than there is between the air and the solid. In this case, a concave meniscus is formed. If the affinity between the air and the solid is greater, then a convex meniscus is formed. |
| Compare vaporization and evaporation. | 334 | Vaporization is when a liquid turns into vapor, evaporation is when the surface of a liquid turns into gas and rises. |

**Summary**:

Liquids have the medium energy level of the three state. They have enough energy for the molecules to break free from their solid lattice but not enough to separate from being in groups of molecules, they move around slowly and randomly in groups, with a moderate level of kinetic energy to power them. Formation of a meniscus is a surface tension effect. Depending on the chemical properties of the various components, there may be a stronger affinity between the liquid and the solid than there is between the air and the solid. In this case, a concave meniscus is formed. If the affinity between the air and the solid is greater, then a convex meniscus is formed. Vaporization is when a liquid turns into vapor, evaporation is when the surface of a liquid turns into gas and rises.